

12) Write the name of the following molecular compounds

CO₂ _____ CO _____

NO₂ _____ N₂O _____

PF₅ _____ P₂O₅ _____

SCl₃ _____ SeI₅ _____

PF₅ _____ P₂O₅ _____

CCl₄ _____ H₂S _____

13) Write the formula of the following molecular compounds

Dihydrogen oxide _____ arsenic pentaiodide _____

Hydrogen chloride (assume all mono) _____ carbon tetrabromide _____

Dinitrogen pentoxide _____ carbon disulfide _____

Sulfur trioxide _____ dinitrogen tetrachloride _____

14) Determine if the following statements refer to **ionic compounds (i)** or **molecular compounds (m)**

ions (anions and cations) _____ molecules _____ sharing electrons _____ gaining and losing e⁻ _____

metal + nonmetal _____ all nonmetals _____ covalent bonds _____ ionic bonds _____

H₂O _____ CO₂ _____ NaCl _____ BaO _____

brittle solid crystals at room temperature _____ electrolytes _____

can be solid or liquid or gas at room temperature _____ usually not electrolytes _____